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Here k_{CP} and k_{PC} are the first order inter-compartmental distribution constants, k_A is the rate constant for absorption from the stomach into the central compartment, k_R is the drug's specific metabolism rate constant and k_e is the elimination rate constant, all with units of reciprocal hours. Oral administration is sometimes considered a ...

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You can tell the overall reaction order by the units of k The activation energy is a "measure" of the energy a that the reacting molecules must have in order for the reaction to occur. The reaction of $AB + C$ to form $A + BC$ is shown above along the reaction coordinate.

Elements of Chemical Reaction Engineering

Parameter Sensitivity Study of the Unreacted-Core Shrinking Model: A Computer Activity for Chemical Reaction Engineering Courses. Journal of Chemical Education 2011 , 88 (1) , 56-58.

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1 Chemical reactions 1.1 Rate of reaction and dependence on temperature We will once again look at the formation of ammonia (NH_3) from nitrogen and hydrogen (see section Chemical equilibrium of the thermodynamics chapter). This reaction follows the equation: $N_2 + 3H_2 \rightleftharpoons 2NH_3$ (1) $H_0 = 92 \text{ kJ mol}^{-1}$ $S_0 = 192 \text{ J mol}^{-1} \text{ K}^{-1}$

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